Physical Science Chapter 6 Test

Completion

Complete each sentence or statement.

1. Garden soil and potato salad are two examples of ________________ mixtures.
2. A type of suspension that does not settle out under normal circumstances is a(n) _________________.
3. A solution is an example of a(n) ________________ mixture.
4. The particles in a colloid are ________________ than those in a suspension.
5. A solution is more ________________ than a suspension.
6. Egg whites, paint, blood, and whipped cream are all examples of _________________.
7. You can separate two ________________ by pouring the less dense liquid off the top.
8. When a solid is dissolved in water, you can separate the two by ________________ or ________________.
9. The process used to separate two miscible liquids is ________________.
10. As sugar dissolves in water, sugar molecules ________________, or spread throughout the entire solution.
11. A solute's ________________ can often be increased by heating.
12. Shaking or stirring a solution will make a solute ________________ more quickly.
13. A solute will dissolve more quickly if you increase its ________________ by breaking it into small pieces.
14. Solutes dissolve faster if the solvent is ________________.
15. Water is a(n) ________________ compound because its shared electrons are not spread evenly throughout each molecule.
16. Because so many substances can dissolve in water, it is often referred to as the ________________.
17. A(n) ________________ solution is an unstable system.
18. A saturated solution contains the greatest quantity of ________________ that will dissolve in a given quantity of ________________.
19. A(n) ________________ is a compound that can change color in a solution, depending on whether the solution is acidic or basic.
20. An acid is a substance that donates hydrogen ions (H⁺) to form ________________ ions when dissolved in water.
21. A(n) ________________ is a substance that either contains hydroxide ions (OH⁻) or reacts with water to form hydroxide ions.
22. Apple juice has a pH of 3, and stomach acid has a pH of 2. This means that stomach acid is ________________ times more acidic than apple juice.
23. Baking soda has a pH of 9, and household ammonia has a pH of 12. This means that ammonia is ______________ times more basic than baking soda.

24. pH is a measure of the ______________ of a solution.

25. In a neutralization reaction, hydronium ions react with hydroxide ions to produce ________________.

26. Salts are ________________ formed when acids and bases react.

27. Salts are ionic compounds that are often soluble in ________________.

28. Because lye contains hydroxide ions, it is a ________________ compound.

29. Baking soda and baking powder are examples of ________________ that are used in cooking.

30. Another name for aspirin is ________________.

31. Soap works because the negatively charged end of the hydrocarbon chain dissolves in ________________, whereas the neutral end dissolves in ________________.

32. Soap is made by reacting compounds made from ________________ with a solution of sodium or potassium hydroxide.

33. A byproduct of the chemical reaction that makes soap is ________________.

34. ________________ make an upset stomach feel better because they ________________ stomach acid.

35. Another name for vitamin C is ________________.

36. Yogurt is made from milk when special bacteria turn ________________, a sugar in milk, to ________________.
Physical Science Chapter 6 Test
Answer Section

COMPLETION

1. heterogeneous
2. colloid
3. homogeneous
4. smaller
5. uniform
6. colloids
7. immiscible liquids
8. evaporation; boiling or filtering
9. distillation
10. diffuse
11. solubility
12. dissolve
13. surface area
14. heated; hot
15. polar
16. universal solvent
17. supersaturated
18. solute; solvent
19. indicator
20. hydronium
21. base
22. 10
23. 1,000
24. hydronium ion concentration
25. water
26. ionic compounds
27. water
28. basic
29. bases
30. acetylsalicylic acid
31. water; oil
32. animal fat
33. glycerol
34. Antacids; neutralize
35. ascorbic acid
36. lactose; lactic acid