Write the complete, ionic and net ionic balanced equations for the reaction of chromium(III) sulfate with barium acetate. Don’t forget to indicate the precipitate.

Complete:
Ionic:
Net Ionic:

Write the complete, ionic and net ionic balanced equations for the reaction of iron(III) acetate with sodium carbonate. Don’t forget to indicate the precipitate.

Complete:
Ionic:
Net Ionic:

Write the (a) Complete, (b) Ionic and (c) Net Ionic balanced equations for the reaction of nickel(II) nitrate with rubidium carbonate. Remember to identify the precipitate.

Write the (a) Complete, (b) Ionic and (c) Net Ionic balanced equations for the reaction of iron (III) chloride with sodium phosphate.

Write the (a) Complete, (b) Ionic and (c) Net Ionic balanced equations for the reaction of Mn(OH)₂ (aq) with HNO₃ (aq).

Write the balanced complete, total ionic and net ionic equations for the reaction of copper (I) sulfate with calcium nitrate.

The net ionic equation for the neutralization of nitric acid with iron (II) hydroxide is:

a. \( 2 \text{HNO}_3 + \text{Fe(OH)}_2 \rightarrow 2 \text{H}_2\text{O} + \text{Fe(NO}_3\text{)}_2 \)

b. \( \text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^- \)

c. \( 2 \text{H}^+ + \text{Fe(OH)}_2 \rightarrow 2 \text{H}_2\text{O} + \text{Fe}^{2+} \)

d. \( \text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O} \)

Write the complete, total ionic and net ionic equation for the reaction of sodium sulfate with barium chloride.

Consider the following balanced molecular reaction

\[ \text{Pb(NO}_3\text{)}_2 (\text{aq}) + \text{Na}_2\text{SO}_4 (\text{aq}) \rightarrow \text{PbSO}_4 (\text{s}) + 2 \text{NaNO}_3 (\text{aq}) \]

and write the spectator ions

Now write as a net ionic reaction leaving out spectator ions.